

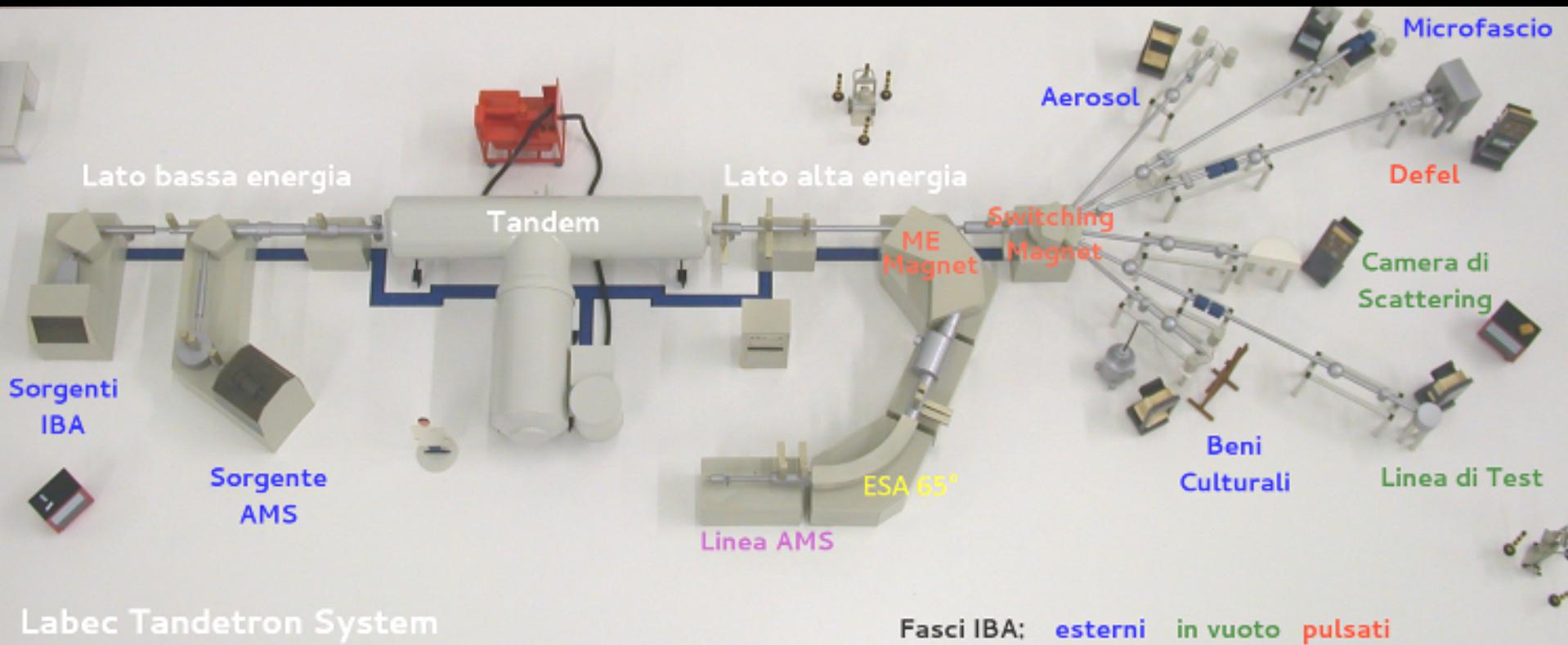


# *Tuning of ion beams and measurements at the LABEC TANDEM*



**Mariaelena Fedi**  
**INFN – Firenze**  
*per il gruppo LABEC Università&Sezione*

# *Measurements at LABEC*



# **Outline**

- ✓ Why measuring by Accelerator Mass Spectrometry (AMS)
  - The example of radiocarbon
- ✓ How to perform an AMS measurement by a Tandem accelerator facility
  - The example of LABEC accelerator
- ✓ **Laboratory session**
  - Beam transport
  - Data analysis



# **Accelerator Mass Spectrometry (AMS)**

- ✓ AMS technique has been introduced to measure the relative abundance of **rare isotopes**
  - Radioactive isotopes of medium-long half life
  - Isotopes characterized by a very low abundance

**$^{14}\text{C}$**   
 $(t_{1/2} = 5730 \text{ yrs})$

**$^{129}\text{I}$**   
 $(t_{1/2} = 1.6 \cdot 10^7 \text{ yrs})$

**$^{10}\text{Be}$**   
 $(t_{1/2} = 1.5 \cdot 10^6 \text{ yrs})$

**$^{36}\text{Cl}$**   
 $(t_{1/2} = 3 \cdot 10^5 \text{ yrs})$



*Archaeometry  
Geology  
Palaeoclimate*



*Environmental  
science  
Nuclear  
safeguards*



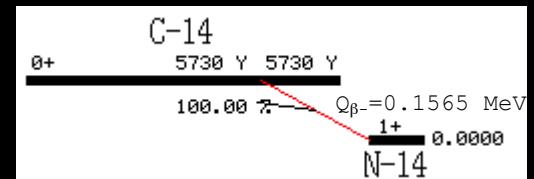
*Environmental  
science  
Geology*



*Environmental  
science  
Geology*

# The example of $^{14}\text{C}$

- ✓ Radiocarbon is the only natural radioactive carbon isotope
- ✓ It is produced in the atmosphere and enters in all the biosphere



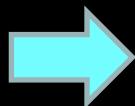
→ **equilibrium condition:**

$$^{14}R = \frac{^{14}\text{C}}{\text{C}_{\text{totale}}} \approx \frac{^{14}\text{C}}{^{12}\text{C}} \approx 1.2 \cdot 10^{-12}$$

- ✓ After the death of an organism, its  $^{14}\text{C}$  content starts to decrease due to the radioactive decay

*Dating of organic findings*

$$^{14}R(t) = ^{14}R_0 e^{-t/\tau}$$



$$t = \tau \cdot \ln\left(\frac{^{14}R_0}{^{14}R(t)}\right)$$

# How to measure rare isotopes

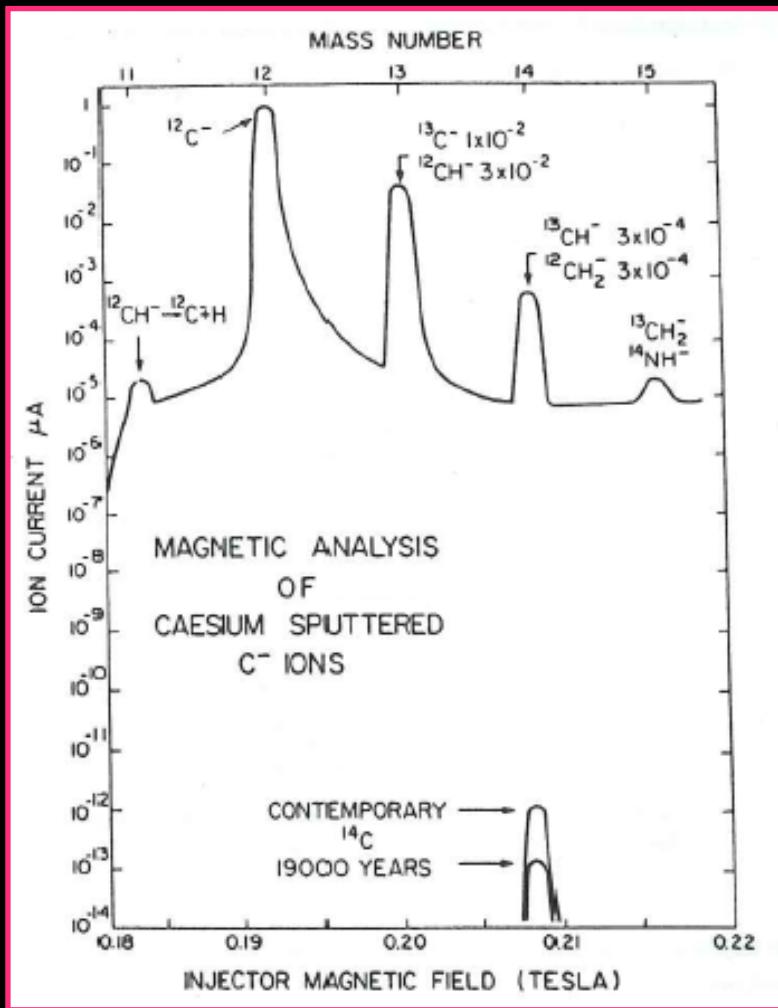
- ✓ By counting the decays
  - however, this cannot be convenient: very low activities → very long measuring time or pretty big samples
- ✓ By counting the atoms
  - however, their discrimination is not easy due to the presence of **isobars**

		120	130	140	150	160	170	180	190
	10N	11N	12N	13N	14N	15N	16N	17N	18N
8C	9C	10C	11C	12C	13C	14C	15C	16C	17C
7B	8B	9B	10B	11B	12B	13B	14B	15B	16B
6Be	7Be	8Be	9Be	10Be	11Be	12Be	13Be	14Be	
4Li	5Li	6Li	7Li	8Li	9Li	10Li	11Li		
3He	4He	5He	6He	7He	8He	9He	10He		
1H	2H	3H	4H	5H	6H				
		1n							

## ISOBARI DEL $^{14}\text{C}$

	$^{14}\text{N}$	$^{12}\text{CH}_2$	$^{13}\text{CH}$
$\Delta m/m$	0.000012	0.0009	0.0003

# *<sup>14</sup>C: a needle in a haystack*



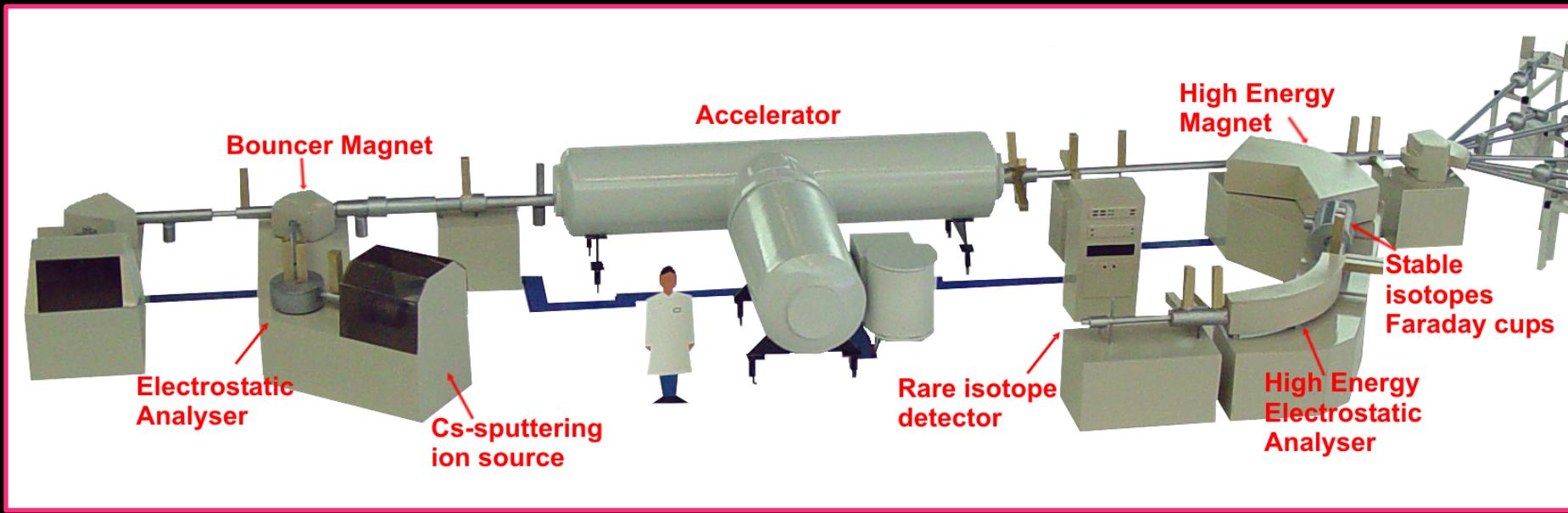
✓ The typical sensitivity of traditional mass spectrometry is not sufficient



**In AMS, electrostatic and magnetic “filters” are coupled to a Tandem electrostatic accelerator**

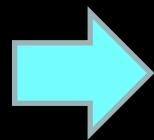
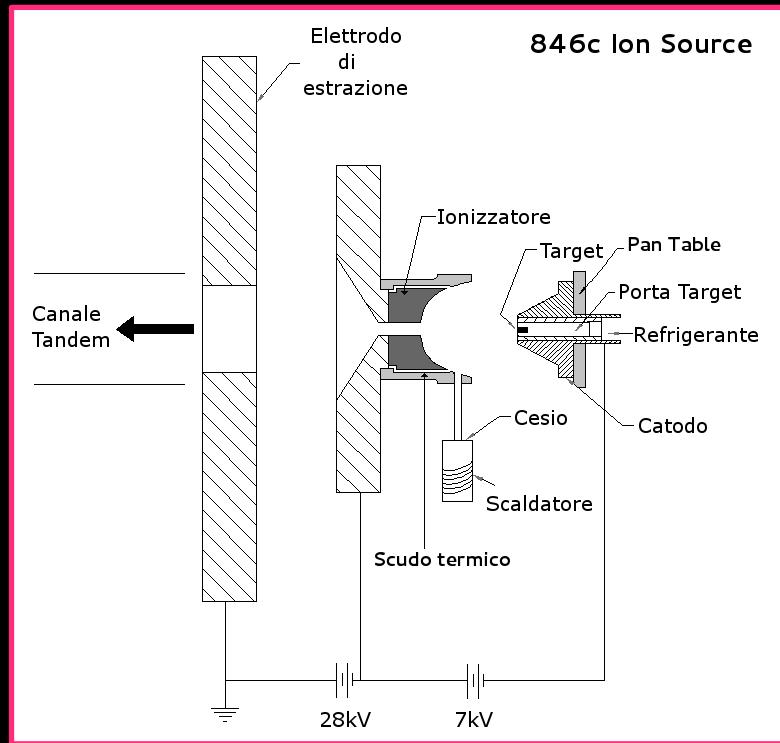
*The accelerator itself is used as a sort of high sensitivity detector*

# *The AMS beam line at LABEC*



# *From the sample to the ions: the source*

- ✓ Samples are inserted in the Cs-sputtering ion source as graphite pellets ( $m < 1 \text{ mg}$ )



**Negative ions are extracted from the source  
→ the isobar  $^{14}\text{N}$  is suppressed**

# Suppression of $^{14}\text{N}$

Ionization Potentials and Electron Affinities of the Elements								
IA	IIA	IIIA	IVA	VA	VIA	VIIA	VIIIA	
$^1\text{H}$ 13.59 0.754								$^2\text{He}$ 24.48 0.078*
				Ionization Potential Electron Affinity				
$^3\text{Li}$ 5.39 0.618	$^4\text{Be}$ 9.32 0.195*	$^5\text{B}$ 8.30 0.277	$^6\text{C}$ 11.26 1.263	$^7\text{N}$ 14.53 -0.07	$^8\text{O}$ 13.61 1.461	$^9\text{F}$ 17.42 3.399	$^{10}\text{Ne}$ 21.56 $< 0$	
$^{11}\text{Na}$ 5.14 0.548	$^{12}\text{Mg}$ 7.64 $< 0$	$^{13}\text{Al}$ 5.98 0.441	$^{14}\text{Si}$ 8.15 1.385	$^{15}\text{P}$ 10.48 0.747	$^{16}\text{S}$ 10.36 2.077	$^{17}\text{Cl}$ 13.01 3.617	$^{18}\text{Ar}$ 15.76 $< 0$	
$^{19}\text{K}$ 4.34 0.501	$^{20}\text{Ca}$ 6.11 0.043	$^{31}\text{Ga}$ 6.00 0.30	$^{32}\text{Ge}$ 7.90 1.2	$^{33}\text{As}$ 9.81 0.81	$^{34}\text{Se}$ 9.75 2.021	$^{35}\text{Br}$ 11.81 3.365	$^{36}\text{Kr}$ 14.00 $< 0$	
$^{37}\text{Rb}$ 4.18 0.486	$^{39}\text{Sr}$ 5.70 $< 0$	$^{49}\text{In}$ 5.79 0.3	$^{50}\text{Sn}$ 7.34 1.2	$^{51}\text{Sb}$ 8.64 1.07	$^{52}\text{Te}$ 9.01 1.971	$^{53}\text{I}$ 10.45 3.059	$^{54}\text{Xe}$ 12.13 $< 0$	
$^{55}\text{Cs}$ 3.89 0.472	$^{56}\text{Ba}$ 5.21 $< 0$	$^{81}\text{Tl}$ 6.11 0.2	$^{82}\text{Pb}$ 7.42 0.364	$^{83}\text{Bi}$ 7.29 0.946	$^{84}\text{Po}$ 8.42 1.9	$^{85}\text{At}$ 9.5 2.8	$^{86}\text{Rn}$ 10.75 $< 0$	

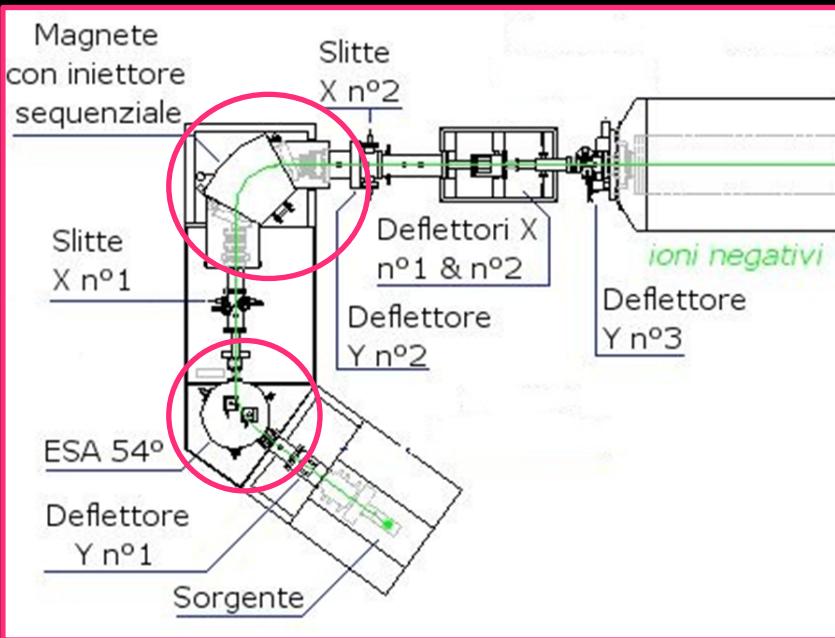
\*Metastable

R. Middleton, A negative-ion cookbook

- ✓  $^{14}\text{N}$  does not form negative ions in the stable state
- ✓ Lifetime of negative excited states are too short with respect to the time needed to travel to the injection into the accelerator
- ✓ Possible fragments from molecules (e.g.NH) are discriminated thanks to the following charge-energy-mass analysis



# The analysis on the low energy side



- ✓ First selection according to **energy/charge**

$$\varepsilon r = 2 \frac{E}{q} \quad q=-1 \quad E= 35 \text{ keV}$$

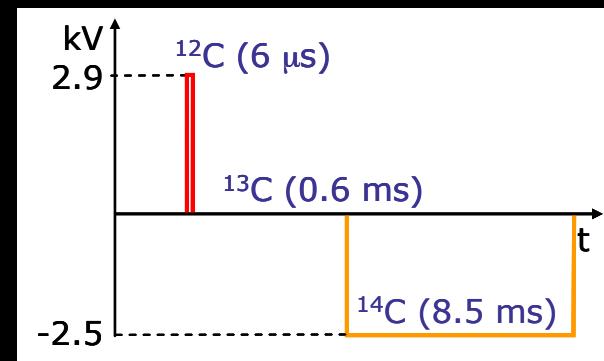
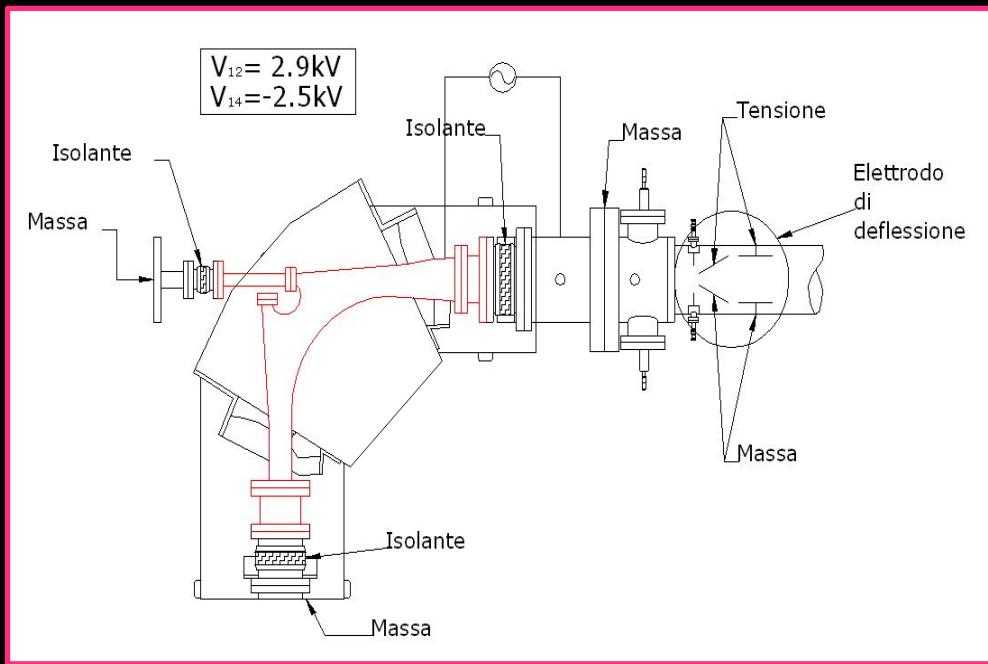
- ✓ Second selection according to **impulse/charge**

$$Br = \frac{\sqrt{2mE}}{q}$$

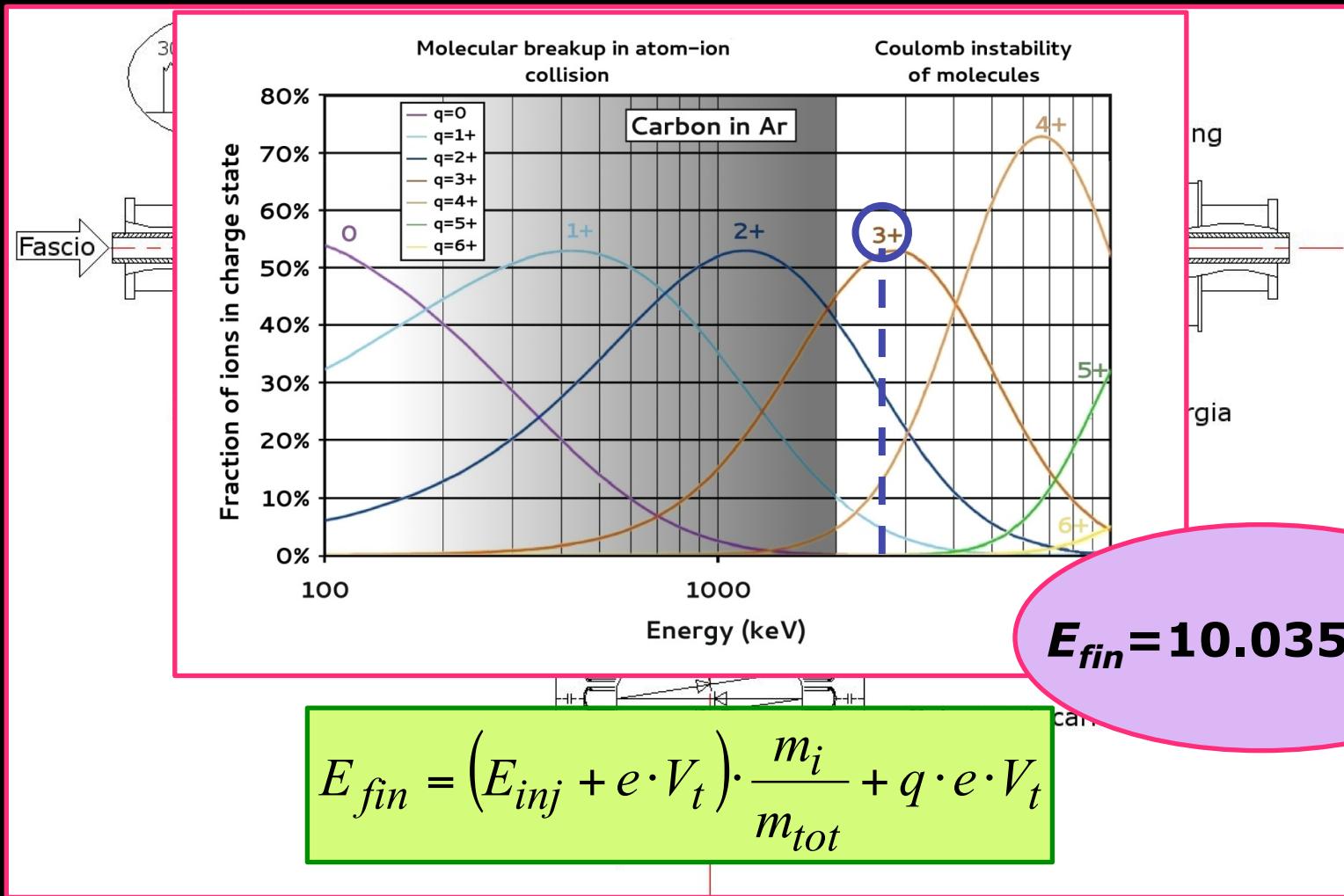
→ to select the three masses 14, 12 and 13

# *Sequential injection...*

- ✓ The magnetic field is kept fixed to optimize the transmission of a mass ( $m=13$ )
- ✓ The other masses are transmitted by changing their energy inside the magnet chamber

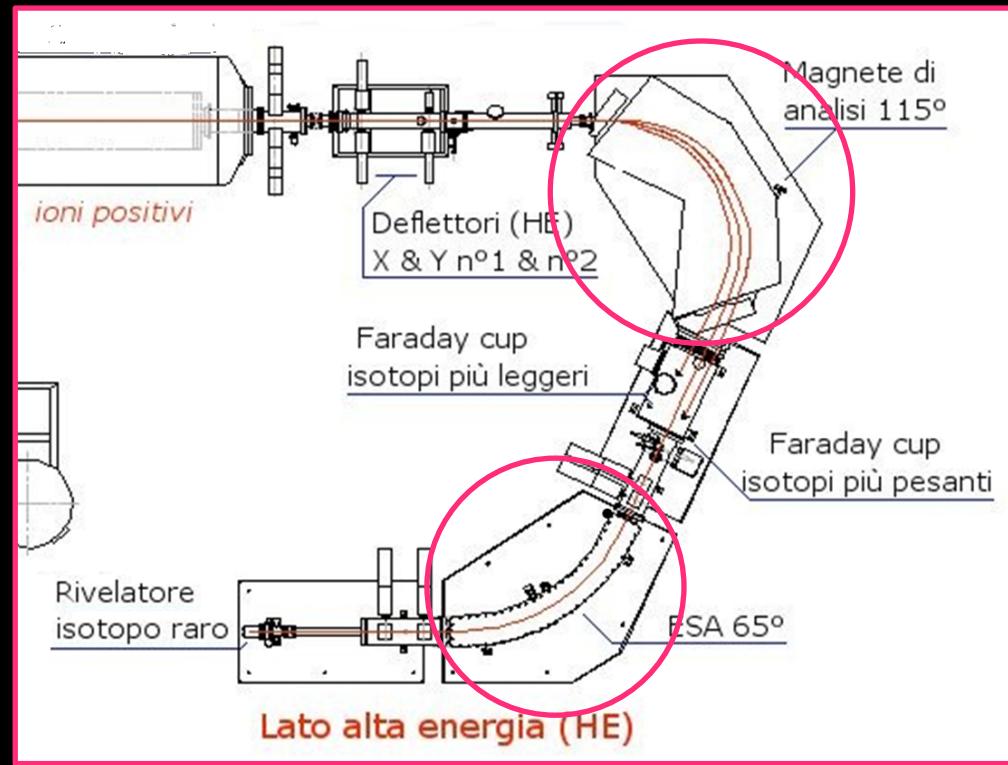


# ...into the accelerator



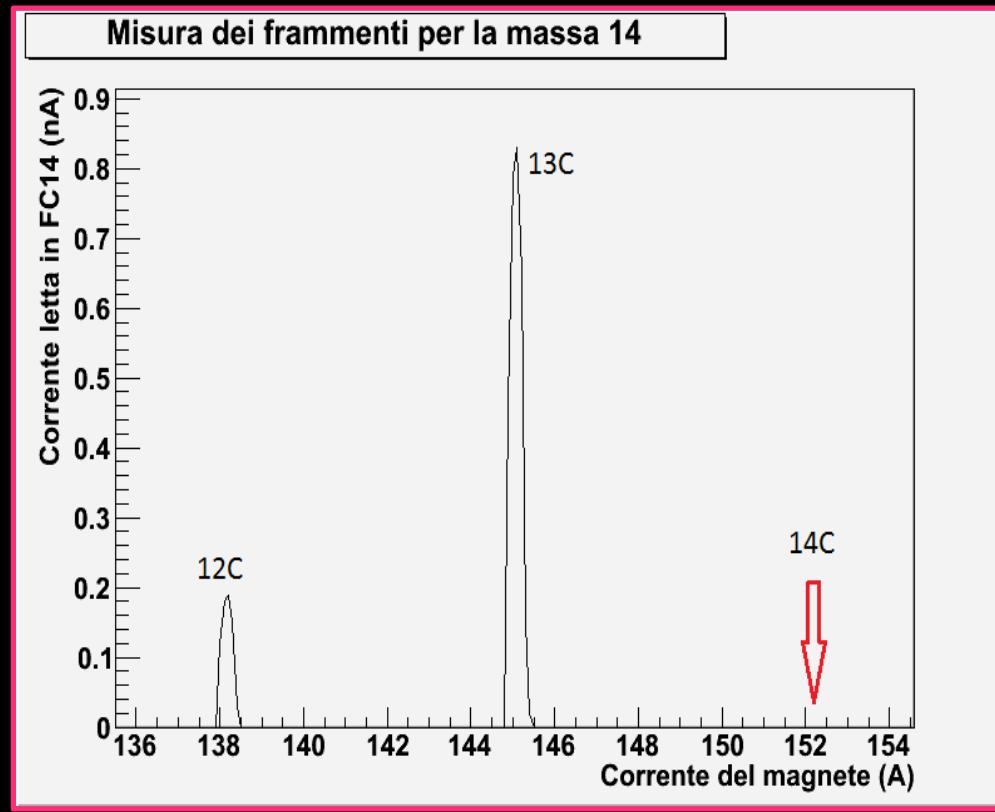
# The analysis on the high energy side

- ✓ First selection according to impulse/charge  
→ **discrimination of  $^{14}\text{C}^{3+}$  ions**
- ✓ Second selection according to energy/charge  
→ rejection of possible residual interferences

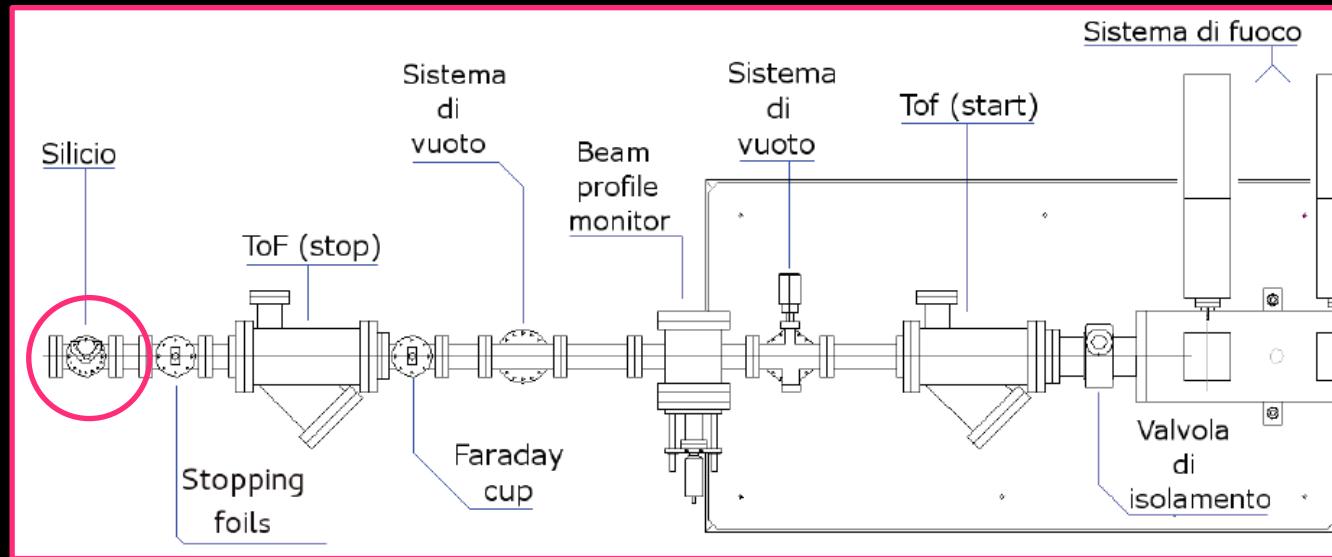


# *What happens to molecular isobars?*

- ✓ As a consequence of the stripping process, molecules can be split into their atomic constituents

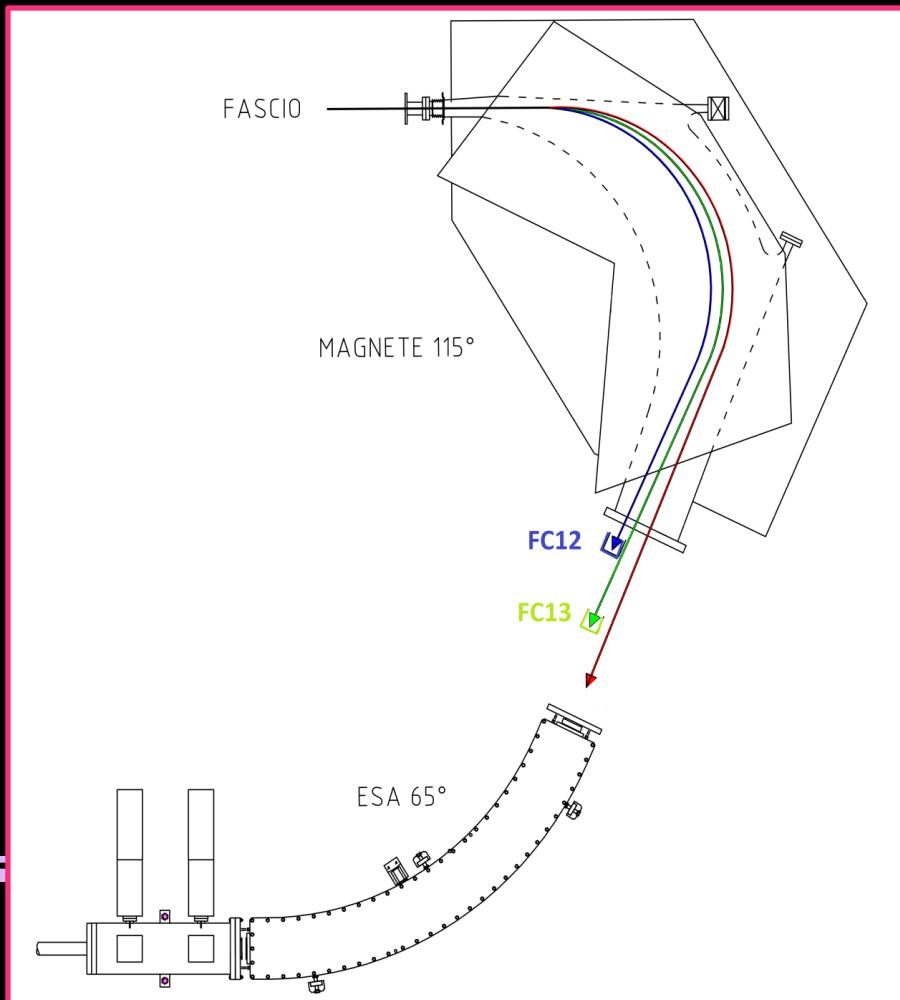


# *Counting the radiocarbon ions*



# What about the stable isotopes?

- ✓  $^{12}\text{C}^{3+}$  and  $^{13}\text{C}^{3+}$  are measured in offset Faraday cups at the exit of the high-energy magnet



Measurement of  $^{12}\text{C}$  →  
measurement of the **isotopic**

$$\text{ratio } \frac{^{12}\text{C}}{^{14}\text{C}}$$

Measurement of  $^{13}\text{C}$  →  
measurement of **isotopic fractionation correction**

# ***The importance of beam transport***

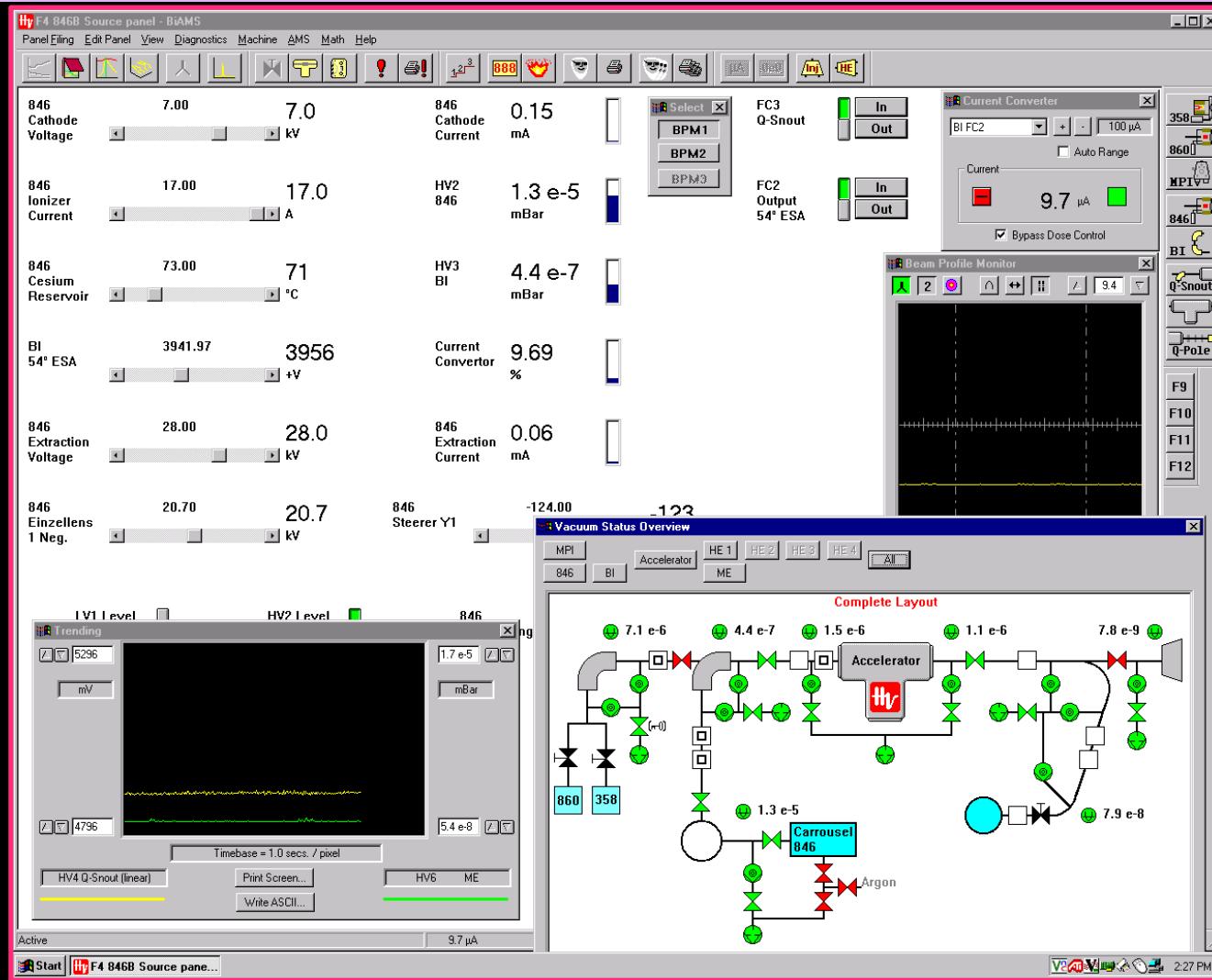
- ✓ The expected  $^{14}\text{C}$  counting rate is low (few tens of Hz at maximum) → the transmission has to be optimized
- ✓ A good analysis of the beam depends on the stability of machine parameters (e.g. energy of the beam, i.e. terminal voltage) → we can exploit feedback systems

***Set focusing elements (Einzel lens and quad doublets) and steerers to optimize the transmission***

***Set electrostatic analysers and magnets to efficiently discriminate the components of the beam***

***Check Faraday cups, Beam Profile Monitors (BPM) and  $^{14}\text{C}$  counting rate to monitor the beam transport***

# Working with the beam



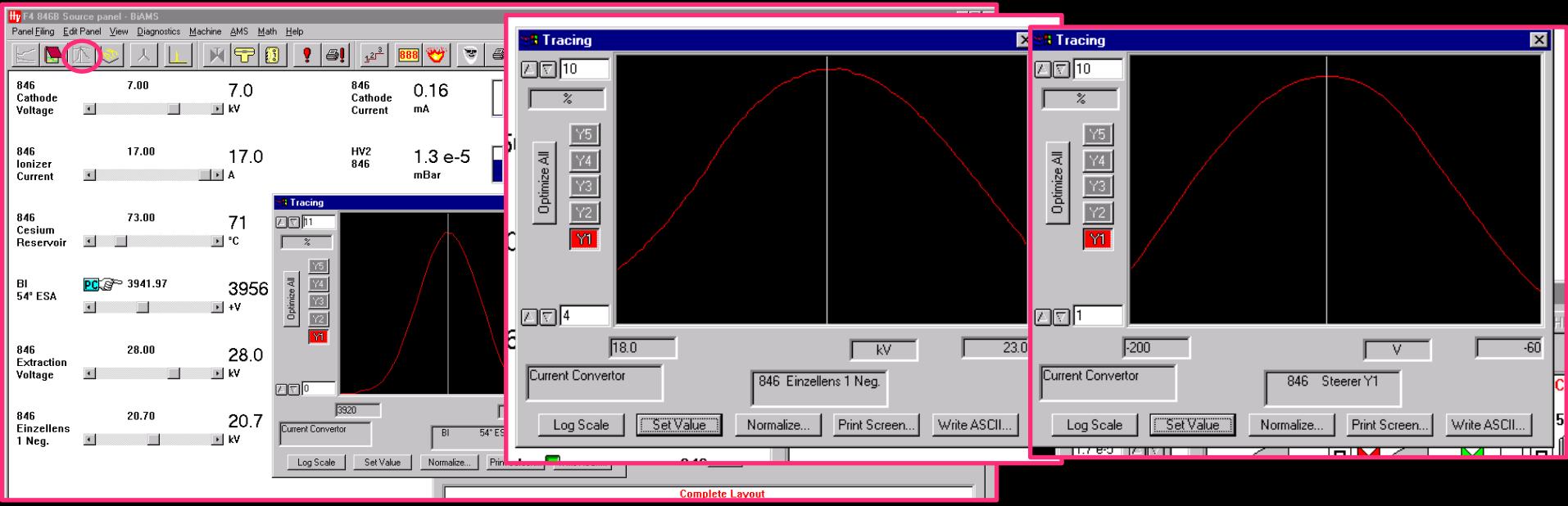
# Beam in FC2

- ✓ Three elements to set

**Electrostatic  
Analyser (ESA)**

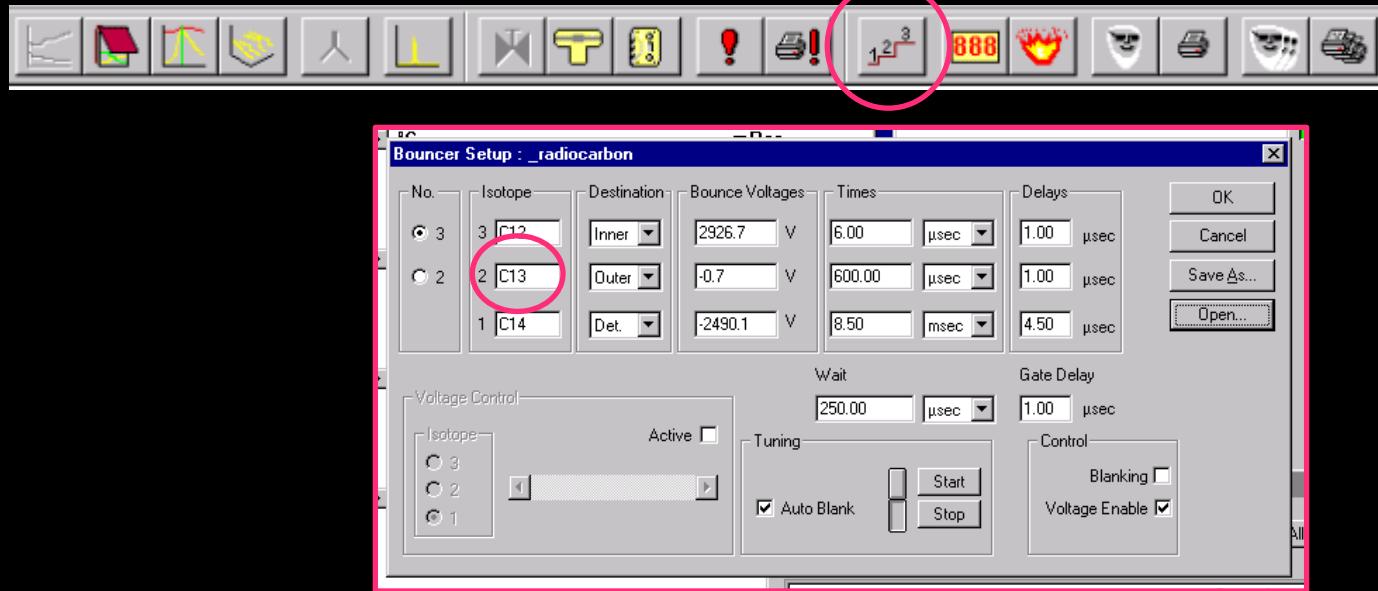
**Einzel Lens**

**Steerer Y**



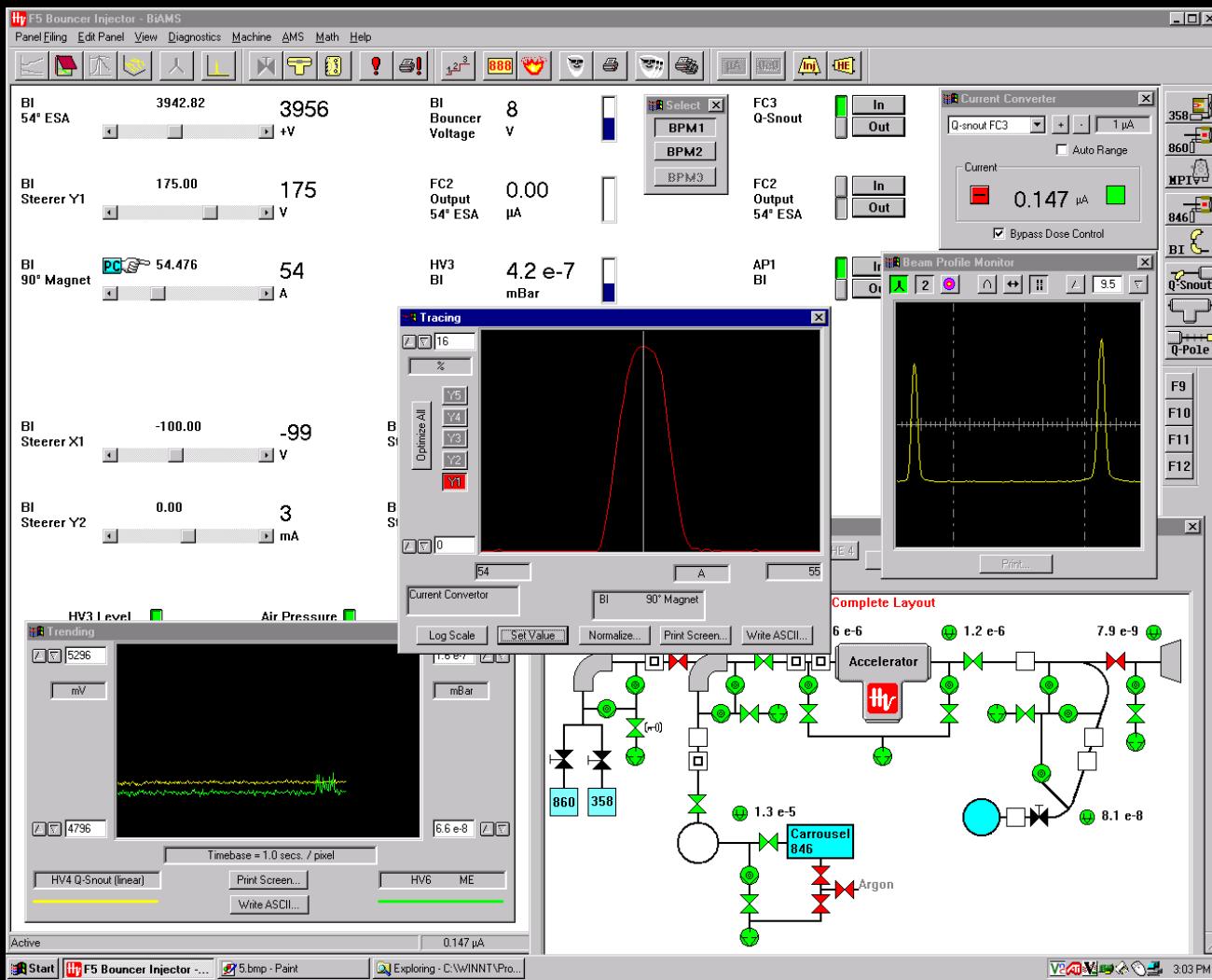
# Beam in FC3

- ✓ Set the magnet → which mass?



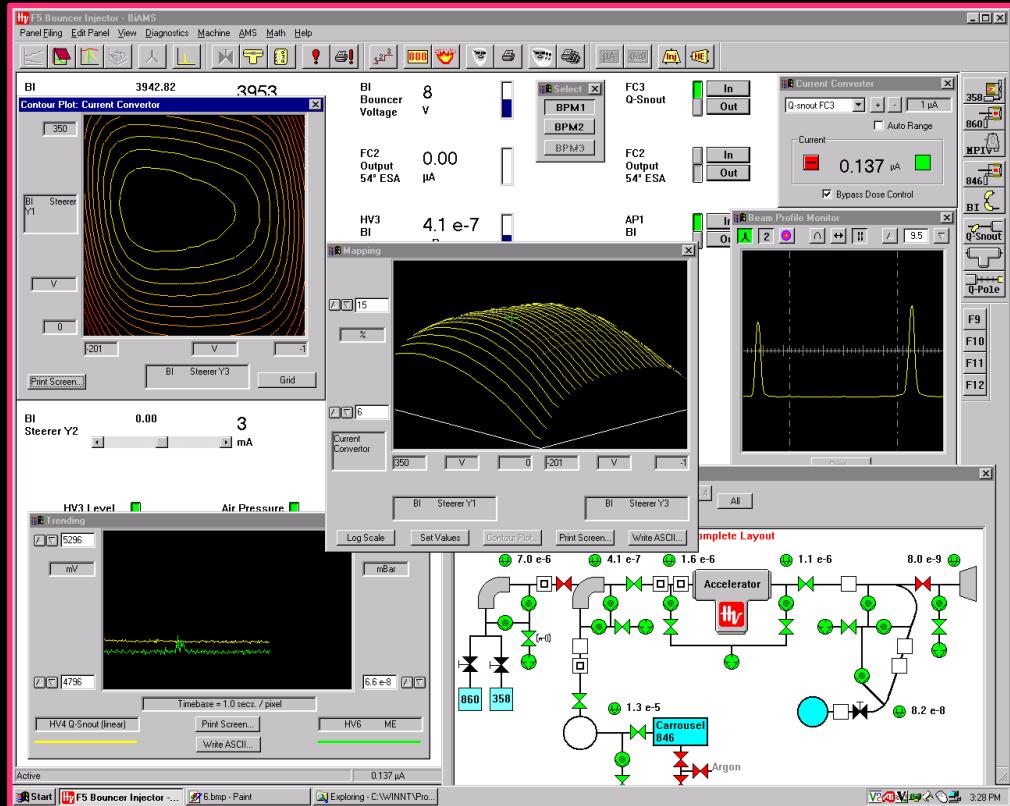
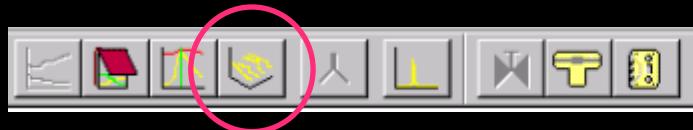
**Optimize the current in the Faraday cup and monitor beam spatial distribution in the BPM**

# Beam in FC3



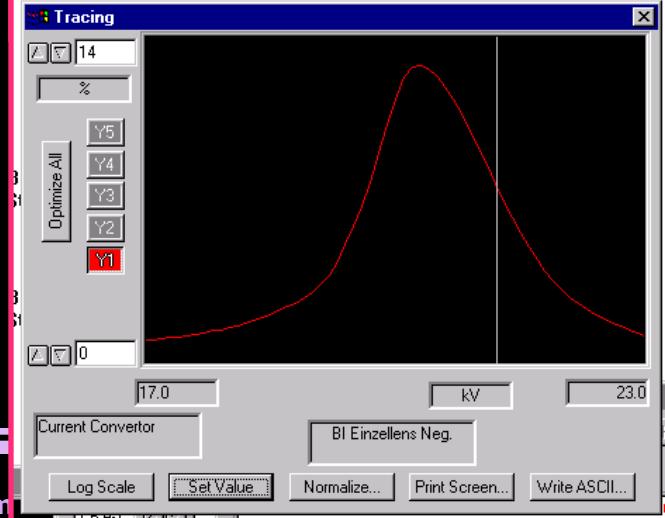
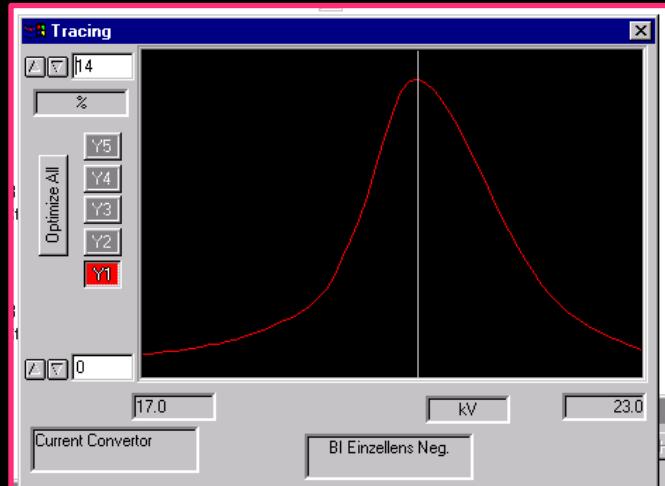
# Beam in FC3

- ✓ A couple of Y steerers
- ✓ A couple of X steerers

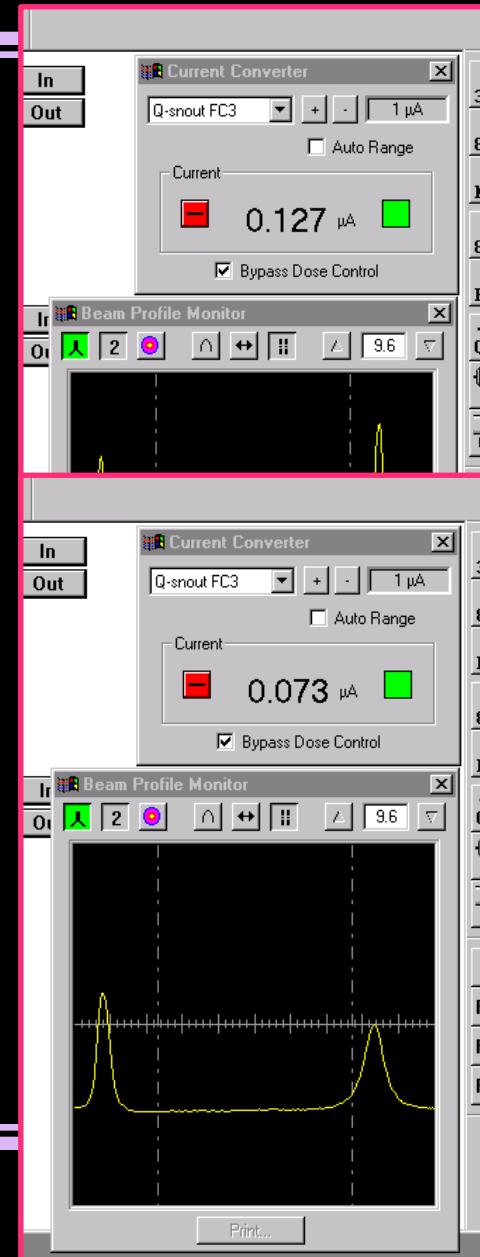
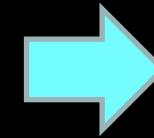
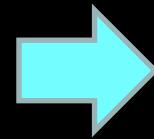


# Beam in FC3

✓ Einzel lens to focus

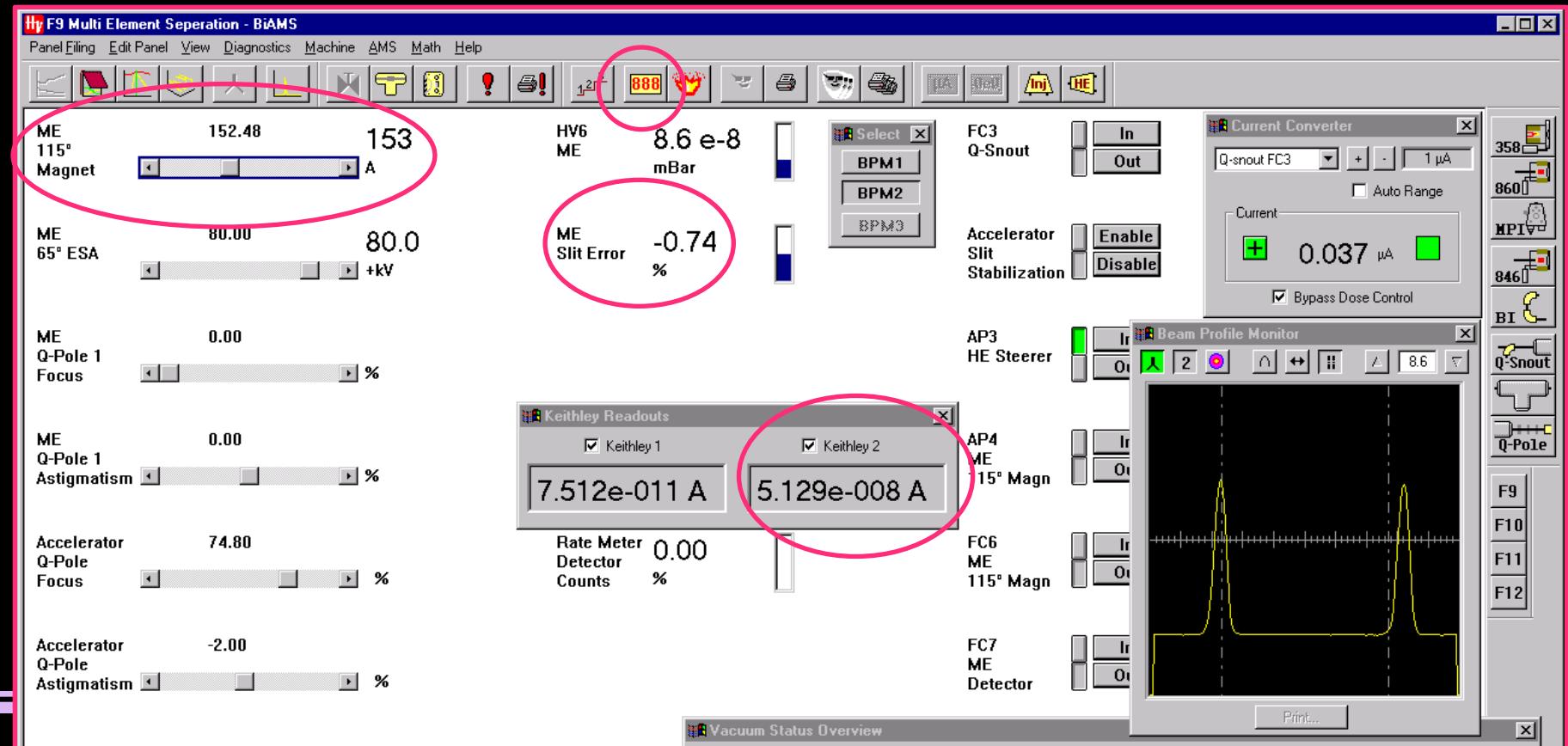


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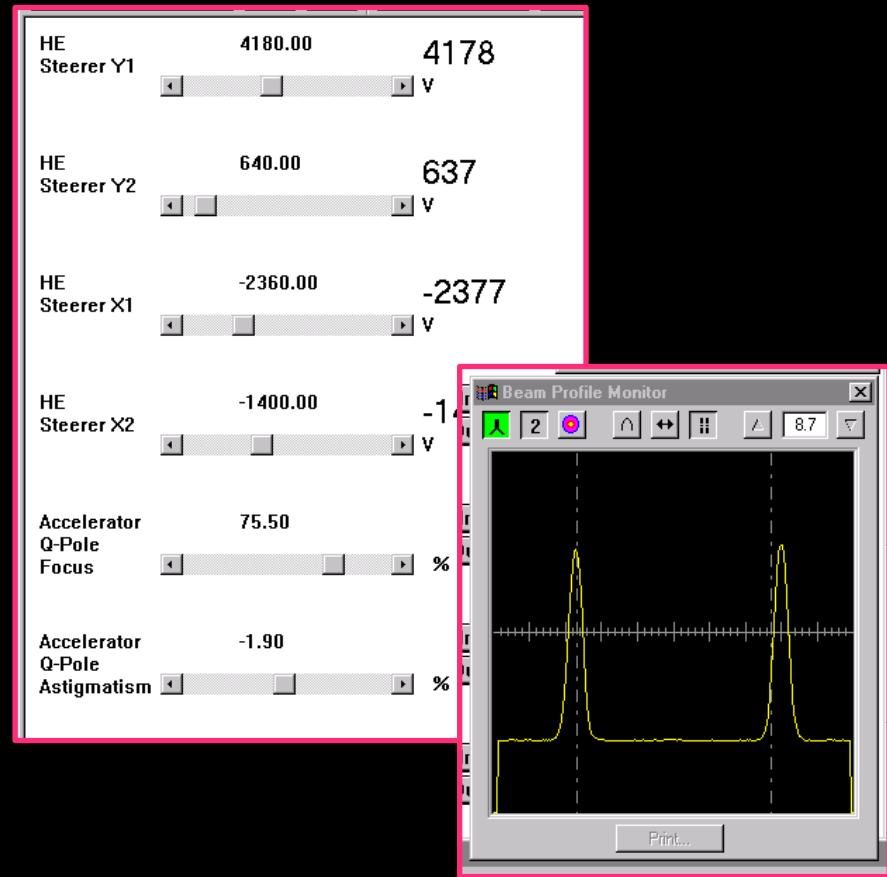
# At the exit of the accelerator

- ✓ Roughly set the magnet on the high energy side → monitor  $^{13}\text{C}^{3+}$  current in the dedicated offset Faraday cup



# *At the exit of the accelerator*

- ✓ Set Q-snout
- ✓ Set quadrupoles  
(focus and astigmatism)  
and high energy  
steerers



# ***Downstream the high-energy magnet***

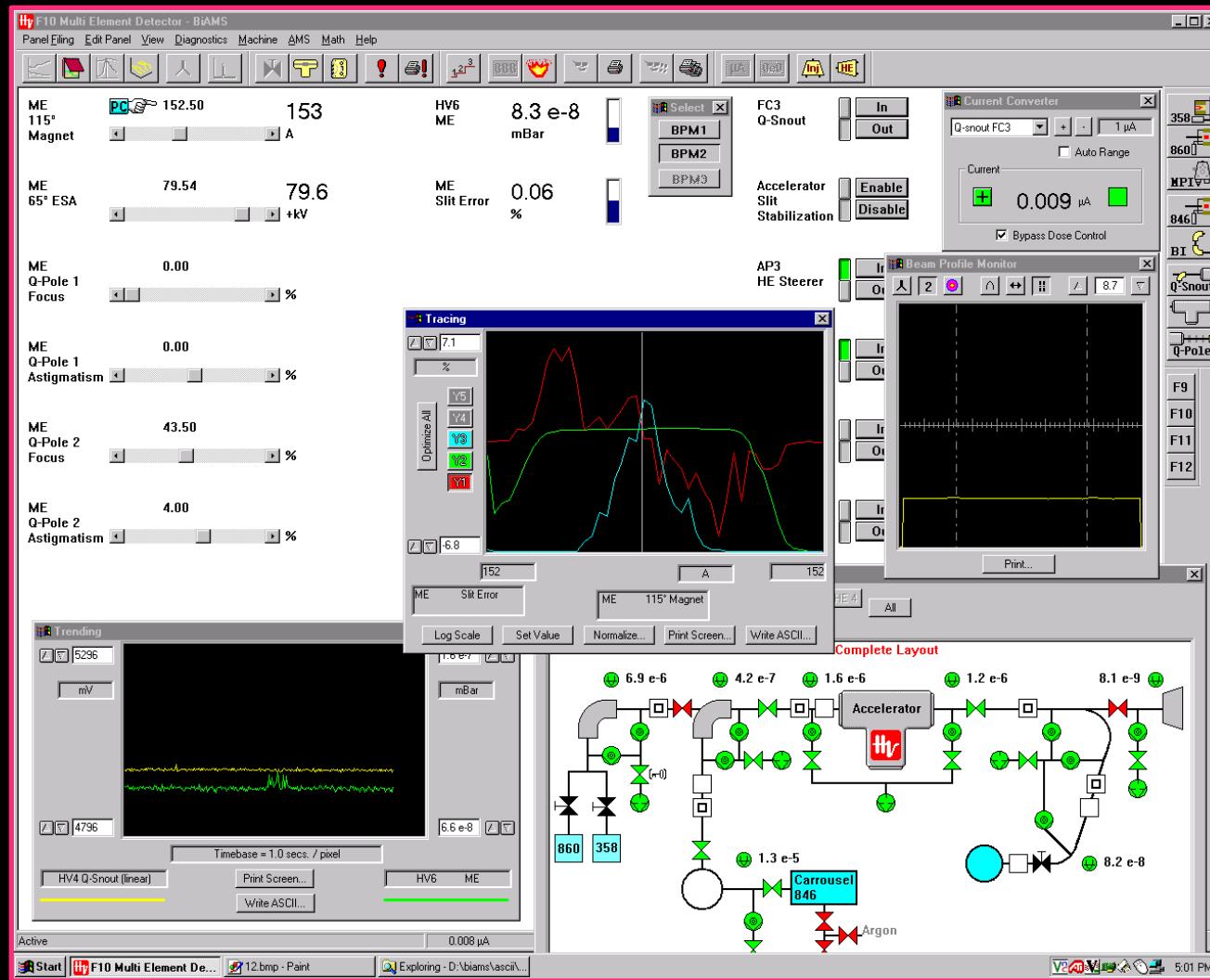
- ✓ Typical beam currents are too low to use Faraday cup or BPM

***Check the counting rate in the  $^{14}\text{C}^{3+}$  silicon detector  
(roughly, looking at the output signal of SCA)***

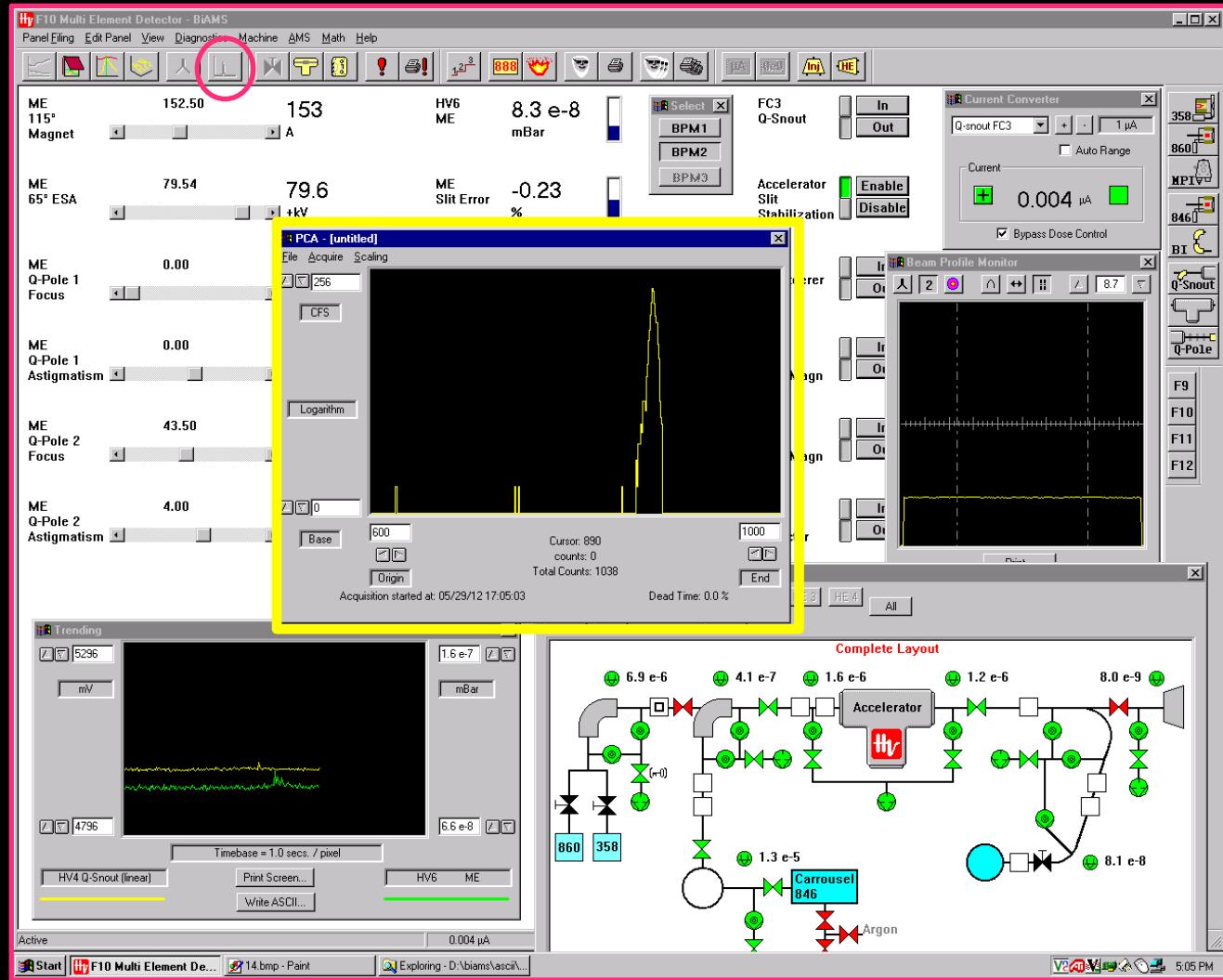
- ✓ Set high-energy electrostatic analyser and quadrupoles



# Final check of the three isotopes



# Our goal



# *The AMS measurement*

- ✓ Your set of samples to be measured:
  - **Standard** samples → to normalize measured isotopic ratio
  - **Blank** samples → to subtract background counts (due to machine and sample preparation)
  - **Unknown samples**
  
- ✓ What is measured:
  - **$^{14}\text{C}$  counts**
  - **$^{12}\text{C}$  and  $^{13}\text{C}$  currents**



# Data Analysis

$$^{14}R_{norm,sample}(pMC) = ^{14}R_{cert,std}(pMC) \cdot \left( \frac{^{13}R_{meas,std}}{^{13}R_{meas,sample}} \right)^2 \cdot \left( \frac{^{14}R_{meas,sample} - ^{14}R_{meas,blk}}{^{14}R_{meas,std} - ^{14}R_{meas,blk}} \right)$$

**Certified concentration of standard**      **Isotopic fractionation correction**      **Normalization to standard and subtraction of background**

$$t_{RC}(\text{yrs BP}) = 8033 \text{ yrs} \cdot \ln \left( \frac{100 \text{ pMC}}{^{14}R_{norm,sample}} \right)$$

**Calibration**

*Most probable calendar age*

# *Lab practice*

<b>14:30 – 15:40</b>	<i>Beam transport and optimization of measurement parameters</i>
<b>15:40 – 15:50</b>	<i>Start of AMS measurement</i>
<b>15:50 – 16:10</b>	<i>Sample preparation laboratory</i>
<b>16:10 – 17:00</b>	<i>Improvements on detection of molecular isobars</i>
<b>17:00 – 18:00</b>	<i>Data analysis</i>

***Have a good job!***